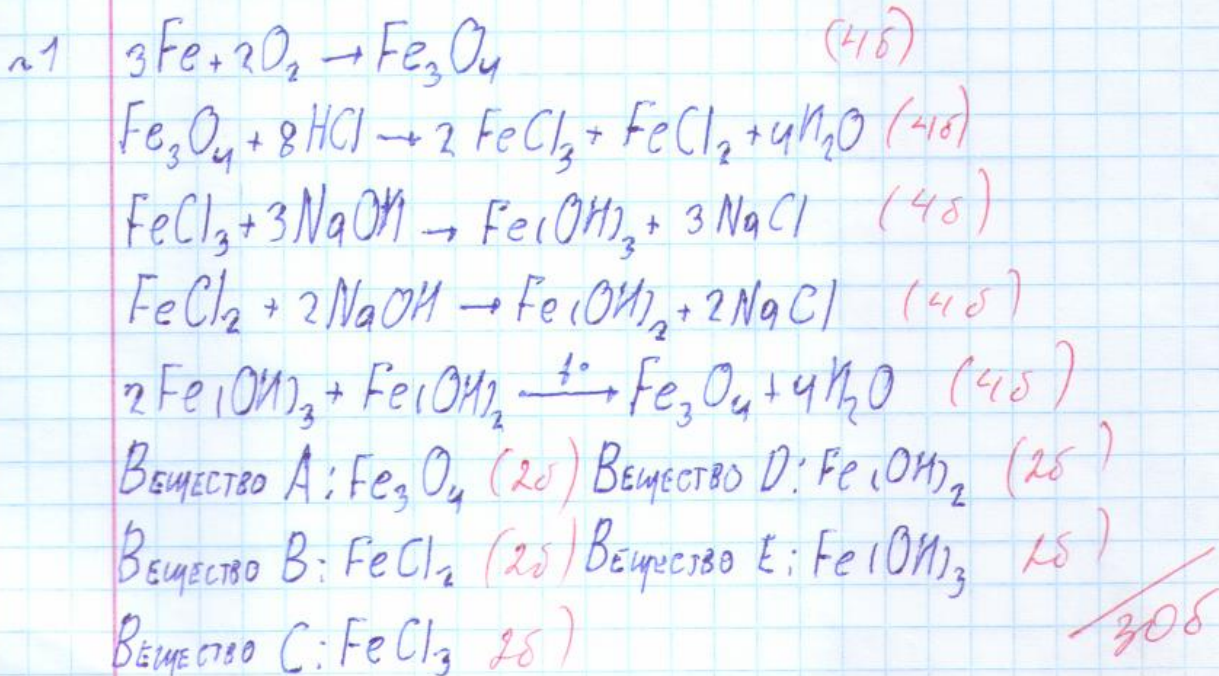




Уаомь 1.

- 1)  $\delta^2 \beta \alpha$  (4)
- 2)  $\beta$  (2)
- 3)  $\alpha$  (2)
- 4)  $\alpha$  (2)
- 5)  $\delta^2 \beta \alpha$  (4)
- 6)  $\beta$  (2)
- 7)  $\alpha$  (4)
- 8)  $\delta$  (2)
- 9)  $\delta$  (2)
- 10)  $\alpha \beta$  (4)  $\swarrow 28$

Уаомь 2.



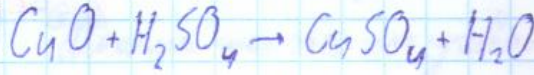
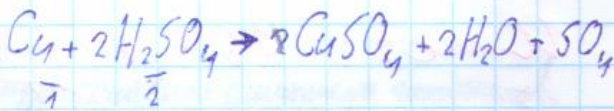
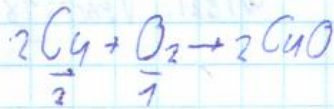
Дано

Решение

$m_{\text{чист}} = 12,96 \text{ г}$

$m_{\text{O}_2} = 2,24 \text{ г}$

$\rho_{\text{р-р (конц)}} = 1,84 \text{ г/мл}$



} 125

$\nu = \frac{m}{M}$

$m = \nu \cdot M$

$V = \frac{m}{\rho}$

1)  $\nu_{\text{O}_2} = \frac{2,24}{32} = 0,07 \text{ моль}$

2)  $\nu_{\text{Cu}} = 0,07 \cdot 2 = 0,14 \text{ моль}$

3)  $m_{\text{Cu}} = 0,14 \cdot 64 = 8,96 \text{ г}$

4)  $m_{\text{CuO}} = m_{\text{чист}} - m_{\text{Cu}}$

$m_{\text{CuO}} = 12,96 - 8,96 = 4 \text{ г}$

5)  $\nu_{\text{CuO}} = \frac{4}{80} = 0,05$

6)  $\nu_1(\text{H}_2\text{SO}_4) = \nu(\text{CuO}) = 0,05$

7)  $\nu_2(\text{H}_2\text{SO}_4) = 2 \cdot \nu_{\text{Cu}}$

$\nu_2(\text{H}_2\text{SO}_4) = \frac{0,14}{0,07} \cdot 2 = 0,28$

8)  $\nu(\text{H}_2\text{SO}_4) = \nu_1(\text{H}_2\text{SO}_4) + \nu_2(\text{H}_2\text{SO}_4)$

$\nu(\text{H}_2\text{SO}_4) = 0,05 + 0,28 = 0,33$

9)  $m(\text{H}_2\text{SO}_4) = 0,33 \cdot 98 = 32,34 \text{ г}$

10)  $m_{\text{раств}} = m(\text{H}_2\text{SO}_4) : 0,98$

$m_{\text{раств}} = \frac{32,34}{0,98} = 33,00 \text{ г}$

} 85

} 125

$$8.5 \quad 11) V_{\text{PACT}} = \frac{19326995}{1,89} = 10,22592391309349$$

$$Q_{\text{TBET}}: V_{\text{PACT}} = 18,3 \text{ мм}$$

400

$$U_{\text{moz}}: 28 + 30 + 40 = 988.$$

Продолжение